

Chemical Quantities Chapter 10

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Chemical Quantities Chapter 10

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Chapter 10 - Chemical Quantities. Jennie L. Borders. Section 10.1 - The Mole: A Measurement of Matter. You often measure the amount of something by count, by mass, or by volume. A mole (mol) of a substance is 6.02×10^{23} representative particles of that substance.

Chapter 10 - Chemical Quantities

CHAPTER 10: Chemical Quantities BASICS: • The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole (mol) of a substance is equivalent to 6.02×10^{23} particles of that substance • The mole was founded by a scientist named Avogadro, and he decided to use the

CHAPTER 10: Chemical Quantities

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Chemical Quantities THE MOLE AND QUANTIFYING MATTER 10.1 The Mole: A Measurement of Matter Essential Understanding The mole represents a large number of very small particles. Reading Strategy Frayer Model The Frayer Model is a vocabulary development tool. The center of the

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Chapter 10 Chemical Quantities Workbook Answers

Guided Notes Chapter 10: Chemical Quantities Name_____ 10.1 The Mole: A Measurement of Matter 1. What is the mole? 2. What is Avogadro's number? 3. What kinds of things can you count using the mole? 4. What are three ways of measuring the amount of a substance? 5.

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Chapter 10 "Chemical Quantities" Vocab Chapter 10 (Chemical Quantities) Test Study Guide The mole is the SI unit used to measure the number of representative particles in a substance. A representative particle can be an atom, an ion, or a molecule, depending upon the way a substance commonly

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Chapter 10 Chemical Quantities Workbook Answers

Chapter 10 "Chemical Quantities" You will need a calculator for this chapter! Calculating Empirical get a ratio from % composition Assume you have a 100 g sample - the percentage become grams ($75.1\% = 75.1$ grams) Convert grams to moles.

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