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Chemistry 12

Worksheet 4 Ka

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**12**

Kb Calculations

Answers

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**Ka Kb**

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\_\_\_\_\_ 37 Due Date

\_\_\_\_\_ Correct and Hand

In by \_\_\_\_\_ NOTE: For

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must show all of your  
steps in each

calculation. State any  
assumptions clearly.

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**37**

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3 of 4 Pages 7.

Calculate the pH of a

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## Chemistry 12

Worksheet 4-Ka

0.22 M solution of the salt  $\text{NaNO}_2$ . Show all of your steps clearly. (6 marks) Answer \_\_\_\_\_ 8.

A 0.40 M solution of the lactate ion ( $\text{C}_3\text{H}_5\text{O}_3^-$ ) (a weak base), has a pH of 8.728.

### **Chemistry 12**

### **Worksheet 4-4 Ka and Kb Calculations**

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Worksheet 4-4. Ka and Kb Calculations. Name \_\_\_\_\_

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Correct and Hand In by \_\_\_\_\_

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NOTE: For this worksheet, you must show all of your steps in each calculation.

State any assumptions clearly. Make sure your answer is in the correct number of significant digits as justified by the data and make sure your answer has the ...

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12-Worksheet 4-4—Ka

& Kb Calculations Page

2 of 4 Pages 3. The pH

in a 0.25 M solution of  
the acid HBrO is 4.65 .

Using this, calculate  
the value of Ka for the  
acid HBrO. (4 marks)

Answer \_\_\_\_\_ 4. The pH  
in a solution of benzoic  
acid is 2.355.

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**Worksheet 4-4Ka Kb**

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Worksheet 4 Ka

Worksheet 4-4—Ka &

Kb Calculations 9. The

weak base ethylamine

( $C_2H_5NH_2$ ) has a  $K_b$  of

$6.4 \times 10^{-4}$  a) Write the

equilibrium equation

for the ionization of

ethylamine. (1 mark)

b) What  $[C_2H_5NH_2]$  is

required to produce an

ethylamine solution

with a pH 12.102? pH:

11.102 / .102% (4 marks)

+0.0126 S 0.0126

0.24 M 0.0065 Co

Answer co

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Worksheet 4-4—Ka Kb

Calculations 3. The pH  
in a 0.25 M solution of  
the acid  $\text{HBrO}$  is 4.65

. Using this, calculate  
the value of  $K_a$  for the  
acid  $\text{HBrO}$ . (4 marks)

4.65 so  $(-4.65) \cdot 10^{-5}$

$3 \times 10^{-5}$  w 4. The

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Worksheet 4 Ka

pH in a solution of benzoic acid is 2.355.

Determine the molar concentration of the benzoic acid. (4 marks)

.00441 + 0.00441b

(0.004416) z

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Neutralization equations (worksheet 4.2), and two types of Ka Calc questions (worksheet 4.3) and Acid-Acid competition questions as they relate to Bronsted-Lowry Conjugate pairs (worksheet 4.4), on Friday Feb 28 (Term 3 mark).

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CHEMISTRY 12  
WORKSHEET 4.10

Titration and Titration  
Curves 1. Calculate the

pH of the solution  
resulting from the  
following reactions: a)

30.0 ml of 0.20 M

NaOH and 30.0 ml of

0.20 M HCl b) 30.0 ml

of 0.20 M NaOH and

40.0 ml of 0.20 M HCl

c) 50.0 ml of 0.10 M

CH<sub>3</sub>COOH and 50.0 ml

of 0.10 M KOH d) 10.0

ml of 0.10 M HClO<sub>4</sub>

and 10.0 ml of 0.10 M

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NH<sub>3</sub> 2.  
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salt solutions.

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show all of your steps  
in each calculation.

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Answers

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Amphiprotic

Calculations. 7. pH

Weak Base

Amphiprotic

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Buffers and Indicators

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Determine the pH of a solution in which 1.00 mol  $\text{H}_2\text{CO}_3$  ( $K_a = 4.2 \times 10^{-7}$ ) and 1.00 mole  $\text{NaHCO}_3$  are dissolved in enough water to form 1.00 L of solution. This is a buffer. The weak acid is carbonic acid and the conjugate base is bicarbonate

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Worksheet 2-3 -

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Calculations Involving  
the Equilibrium

Constant Page 8 has a  
0.400. Exactly 1.00 mol  
of each gas was in a  
100.01. vessel and the  
E O.  $1.363 \times 10^{-3}$   
 $2.932 \times 10^{-2}$  has a  $1.3 \times 10^{-2}$   
Y are injected into a  
1.0 L container at 250C  
, find 14. The reaction.  
Chemistry 12 Unit 2 -  
Chemical Equilibrium

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